

## Ram Seshadri MRL 2031, x6129

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### Submit to Justin by October 14

1. Silicon has the diamond structure (8 atoms per unit cell), a cell parameter at room temperature of  $a = 5.4309 \text{ \AA}$ , and a density of  $2330 \text{ kg m}^{-3}$ . Use this information to estimate the Avogadro number. The atomic mass of Si is 28.086.
2. The cell parameter of bcc Fe is  $a = 2.8665 \text{ \AA}$ . What would the cell parameter be if the structure were fcc instead of bcc (Hint: The atomic radius of Fe could be expected to remain the same.)
3. The cell parameter of NaCl is  $a = 5.63 \text{ \AA}$ . What is its density? What would the density be if 5% of the Na atoms in NaCl are missing? <sup>1</sup>. Remember that there are 4 Na and 4 Cl atoms in the unit cell.
4. We have been projecting structures by considering sections along the  $z$  axis. But we could easily consider other projections:
  - (a) Sketch sections of a single fcc unit cell along the body diagonal [the line connecting the cartesian points (0,0,0) and (1,1,1); the sections are perpendicular to this line].
  - (b) Sketch sections of the CsCl cell along a face diagonal [Sections perpendicular to the line joining (0,0,0) and (1,1,0).]
  - (c) Sketch sections of the NaCl cell along a body diagonal
5. Show using side-by-side projections that the zinc blende structure and the diamond structure are closely related. What does this say about the structures of the following compounds, all of which have the same number of valence electrons per atom as Si: AlP, GaAs and InSb?
6. The diamond structure has very poor packing. Suggest why some elements and compounds might adopt this structure, rather than a structure with higher packing efficiency.

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<sup>1</sup>Such missing atoms give rise to point defects that we will discuss presently