



Filling Order.

Remember:

1. Electrons are filled closest to the nucleus first: This is the *aufbau* principle.
2. s can take 2 electrons, p can take 6, d can take 10, and f can take 14. This is a consequence of the Pauli exclusion principle.
3. 4s fills before 3d, but 4s is ionized *before* 3d
4. d⁵ and d¹⁰ have added stability, and are sometimes preferred. Look at Cr and Cu !